

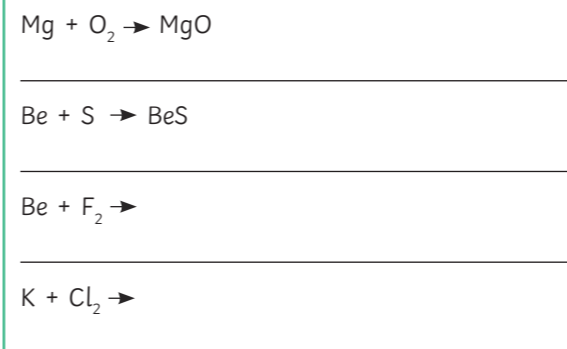
Draw and label an atom. Include labels for the following: neutron, proton, electron.

True or false?  
 1. The radius of an atom is 0.1nm  
 2. Most of the mass is in the shell of the atom.

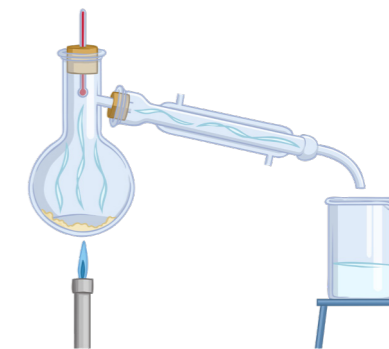
What are the symbols for the following elements?

Element	Symbol
oxygen	
lithium	
sodium	
potassium	
helium	
carbon	
magnesium	

Complete and balance the following equations. What is the name of the compound formed?



**Separating Mixtures**  
 What are the following separation techniques?



\_\_\_\_\_

Fill in the table to show the charges and mass of the components of an atom.

Name	Charge	Relative Mass
proton		
neutron		
electron		

What is the overall charge of an atom?  
 Positive  
 Negative  
 No charge

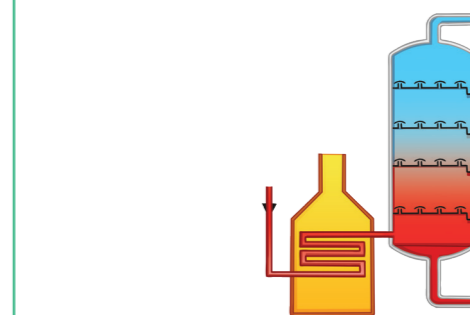
Complete the following diagram for sodium, include the atomic number and the atomic mass number.

Na

What is the mass number?  
 How do you calculate neutron number?

**Mixtures**  
 Write the definition of a mixture. Give two examples.

\_\_\_\_\_  
 \_\_\_\_\_  
 \_\_\_\_\_  
 \_\_\_\_\_



\_\_\_\_\_

Name the compounds and the elements they contain.

NaCl - \_\_\_\_\_  
 MgO - \_\_\_\_\_  
 MgS - \_\_\_\_\_  
 FeS - \_\_\_\_\_

A compound is 2 or more \_\_\_\_\_, chemically \_\_\_\_\_.

Which of the following are compounds?  
 Put a ring round them.  
 oxygen, salt water, magnesium oxide, sodium chloride, nitrogen

Why have you circled the ones you have?  
 \_\_\_\_\_  
 \_\_\_\_\_

**Isotopes** are elements with a different number of \_\_\_\_\_ but the same number of \_\_\_\_\_, e.g. carbon 12 and carbon 14.

How can you use isotopes to calculate the relative atomic mass? Write down the equation.

What is the ratio of the elements in the following compounds?

e.g. CaO = 1:1      NaCl = \_\_\_\_\_  
 MgCl<sub>2</sub> = \_\_\_\_\_      lithium fluoride = \_\_\_\_\_  
 K<sub>2</sub>O = \_\_\_\_\_      sodium hydroxide = \_\_\_\_\_

What separation technique would you use to separate out different inks in pens?  
 \_\_\_\_\_

How can salt be collected using the process of crystallisation?  
 \_\_\_\_\_  
 \_\_\_\_\_

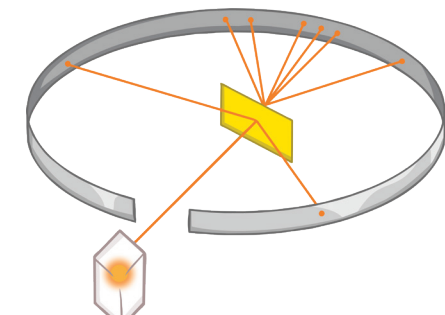
Sand and water can be separated by using a process called \_\_\_\_\_  
 Describe, in 4 steps, how to collect salt from rock salt.

- \_\_\_\_\_
- \_\_\_\_\_
- \_\_\_\_\_
- \_\_\_\_\_

Complete the electronic structure diagrams for:  
**oxygen**  
  
**magnesium**

**a** List 3 halogens  
 \_\_\_\_\_, \_\_\_\_\_, \_\_\_\_\_  
 How many electrons do they have in their outer shell?  
 \_\_\_\_\_  
 Describe how the reactivity changes as you go down the group.  
 \_\_\_\_\_  
 \_\_\_\_\_  
 \_\_\_\_\_  
**d** Complete the following dot and cross diagrams for:  
**NaCl**  
  
**MgO**  
 Write balanced symbol equations for the following reactions:  
 bromine + potassium iodide  
 chlorine + sodium iodide  
 fluorine + potassium chloride

**g** Describe the plum pudding model of the atom.  
 Draw a diagram.  
 \_\_\_\_\_  
 \_\_\_\_\_  
 Why did scientists believe this model?  
 \_\_\_\_\_

**j** Describe what the alpha scattering experiment showed scientists.  
  
 \_\_\_\_\_  
 \_\_\_\_\_

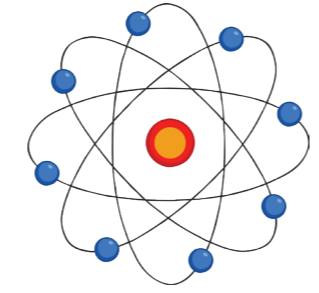
**b** Describe why the noble gases are so unreactive.  
 \_\_\_\_\_  
 \_\_\_\_\_  
 \_\_\_\_\_  
 The boiling points of the noble gases increase/decrease as you go down the group. (delete the wrong answer) Can you explain your answer?  
 \_\_\_\_\_  
 \_\_\_\_\_  
 \_\_\_\_\_

**e** Underline the properties of metals and circle the properties of non-metals:  
 Strong, low density, malleable, dull, good conductors of heat and electricity, high melting and boiling point, brittle, not good conductors of electricity.

**h** Complete word equations for the following reactions:  
 sodium + chlorine →  
 lithium + iodine →  
 potassium + bromine →

**k** Describe what happens to the reactivity of the alkali metals as you go down the group.  
 \_\_\_\_\_  
 Why?  
 \_\_\_\_\_  
 \_\_\_\_\_  
 Complete the word and symbol equation for sodium reacting with water:  
 sodium + water → sodium hydroxide + \_\_\_\_\_  
 Na + \_\_\_\_\_ → NaOH + \_\_\_\_\_

**c** Describe what happens to the reactivity of the alkali metals as you go down the group.  
 \_\_\_\_\_  
 Why?  
 \_\_\_\_\_  
 \_\_\_\_\_  
 Complete the word and symbol equation for sodium reacting with water:  
 sodium + water → sodium hydroxide + \_\_\_\_\_  
 Na + \_\_\_\_\_ → NaOH + \_\_\_\_\_

**f** James Chadwick discovered the... (underline the correct answer)  
 proton  
 neutron  
 electron  


**i** How are the groups arranged in the periodic table?  
 \_\_\_\_\_  
 \_\_\_\_\_  
 How can you tell that the alkali metals are very reactive?  
 \_\_\_\_\_  
 \_\_\_\_\_  
 How can you tell the noble gases are unreactive?  
 \_\_\_\_\_  
 \_\_\_\_\_

**l** Niels Bohr discovered that  
 \_\_\_\_\_  
 \_\_\_\_\_  
 Why did Mendeleev leave gaps in the periodic table?  
 \_\_\_\_\_  
 \_\_\_\_\_  
 What happened to some of the gaps he left?  
 \_\_\_\_\_  
 \_\_\_\_\_

The **transition** elements are a group of metals with similar properties which are different to the metals in group 1. Shade in the transition metals on the periodic table below.

Name three common transition metals.

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Complete the sentences below to describe the properties of transition metals.

They form \_\_\_\_\_ compounds when reacting.

They are \_\_\_\_\_ of heat and electricity. They are malleable. They have \_\_\_\_\_ densities.

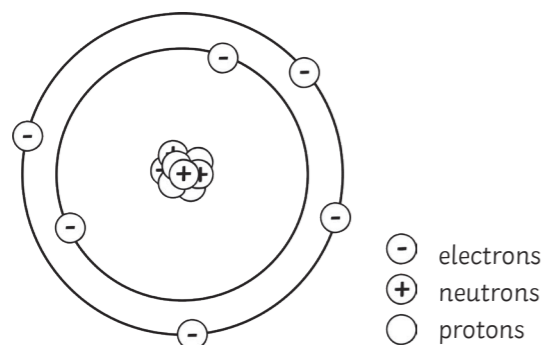
Compared to the alkali metals, they are \_\_\_\_\_ reactive.

Complete the table to show the ions and colours formed by iron compounds. iron (III) oxide

Compound Name	Ion	Colour
iron (II) hydroxide	$\text{Fe}^{2+}$	_____
_____	$\text{Fe}^{3+}$	_____
iron (III) oxide	_____	_____



Draw and label an atom. Include labels for the following: neutron, proton, electron.



⊖ electrons  
⊕ neutrons  
○ protons

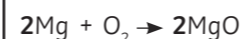
True or false?

- The radius of an atom is 0.1nm **True**
- Most of the mass is in the shell of the atom. **False, most of the mass is in the centre**

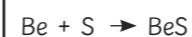
What are the symbols for the following elements?

Element	Symbol
oxygen	O
lithium	Li
sodium	Na
potassium	K
helium	He
carbon	C
magnesium	Mg

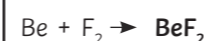
Complete and balance the following equations. What is the name of the compound formed?



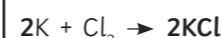
**magnesium oxide**



**beryllium sulphide**



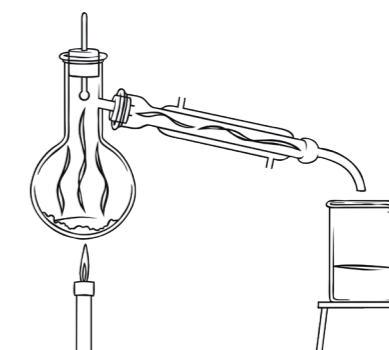
**beryllium fluoride**



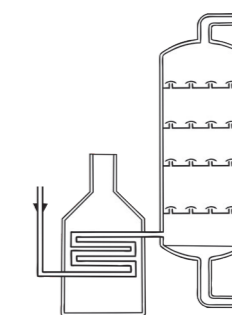
**potassium chloride**

**Separating Mixtures**

What are the following separation techniques?



**Distillation**



**Fractional distillation**

Fill in the table to show the charges and mass of the components of an atom.

Name	Charge	Relative Mass
proton	+1	1
neutron	0	1
electron	-1	very small

What is the overall charge of an atom?

**No charge**

Complete the following diagram for sodium, include the atomic number and the atomic mass number.

23 mass number

**Na**

11 atomic number

What is the mass number?

**Total number of protons and neutrons.**

How do you calculate neutron number?

**Atomic mass – proton number**

A compound is 2 or more **elements**, chemically **joined**.

Which of the following are compounds?

Put a ring round them.

oxygen, salt water, **magnesium oxide**, **sodium chloride**, nitrogen

Why have you circled the ones you have?

**They have 2 or more elements in the word equation.**

**Isotopes** are elements with a different number of **neutrons** but the same number of **protons**, e.g. carbon 12 and carbon 14.

How can you use isotopes to calculate the relative atomic mass? Write down the equation.

$$A_r = \frac{\text{sum of (isotope abundance} \times \text{isotope mass number)}}{\text{sum of abundances of all the isotopes}}$$

**Mixtures**

Write the definition of a mixture. Give two examples.

**Two or more elements together, not chemically joined and can be easily separated.**

**Salt water, sand and water**

Name the compounds and the elements they contain.

NaCl - **sodium chloride, sodium and chlorine**

MgO - **magnesium oxide, magnesium and oxygen**

MgS - **magnesium sulfide, magnesium and sulfur**

FeS - **iron sulfide, iron and sulfur**

What is the ratio of the elements in the following compounds?

e.g. CaO = 1:1

NaCl = 1:1

MgCl<sub>2</sub> = 1:2

lithium fluoride = 1:1

K<sub>2</sub>O = 2:1

sodium hydroxide = 1:1:1

What separation technique would you use to separate out different inks in pens?

**Chromatography**

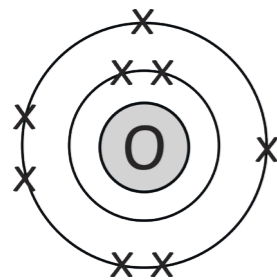
How can salt be collected using the process of crystallisation?  
**By heating up a mixture of salt and water, the water will evaporate and leave the salt in the bowl.**

Sand and water can be separated by using a process called **filtration**.

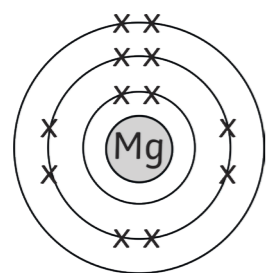
Describe, in 4 steps, how to collect salt from rock salt.

- Grind the mixture;**
- Add water and stir;**
- Filter the mixture;**
- Evaporate the salt water and salt is left over.**

Complete the electronic structure diagrams for:  
oxygen



magnesium



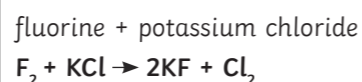
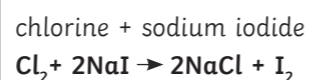
List 3 halogens  
chlorine, fluorine, iodine, astatine

How many electrons do they have in their outer shell?  
7 electrons

Describe how the reactivity changes as you go down the group.

**They become less reactive, the atom becomes larger because there are more electron shells, further from the nucleus so the pull of the nucleus is less. So the electron is less likely to be gained as there is less of a positive pull.**

Write balanced symbol equations for the following reactions:



Describe why the noble gases are so unreactive.

**Their outer shell is full of electrons.**

The boiling points of the noble gases **increase** as you go down the group.

**This is because there are more forces to bond the atoms together, therefore more energy is required to break the bonds.**

Describe what happens to the reactivity of the alkali metals as you go down the group.

**It increases**

Why?

**The number of electrons increases. They are further away from the nucleus. There is less pull on the outer electrons so the atom is more likely to lose an electron.**

Complete the word and symbol equation for sodium reacting with water:



Underline the properties of metals and circle the properties of non-metals:

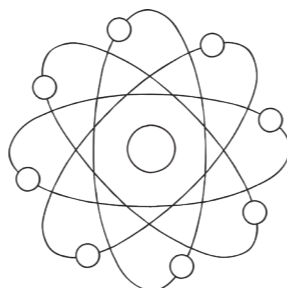
Strong, low density, malleable, dull, good conductors of heat and electricity, high melting and boiling point, brittle, not good conductors of electricity

James Chadwick discovered the...  
(underline the correct answer)

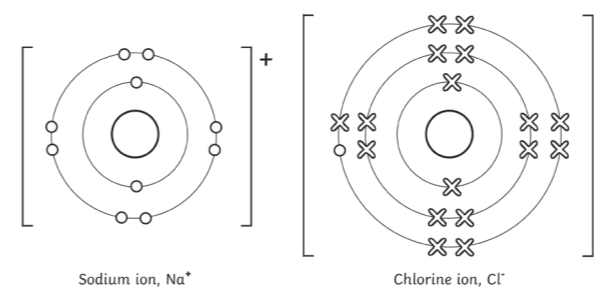
proton

neutron

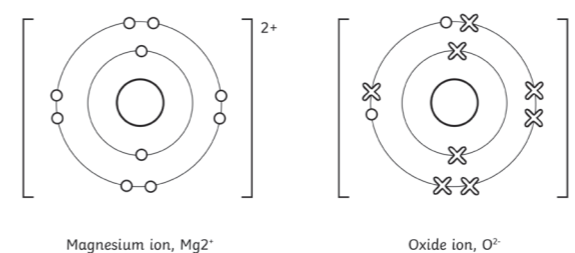
electron



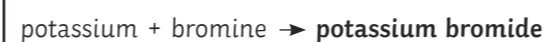
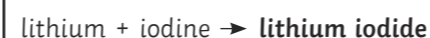
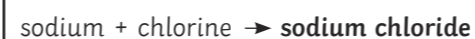
Complete the following dot and cross diagrams for:  
NaCl



MgO



Complete word equations for the following reactions:



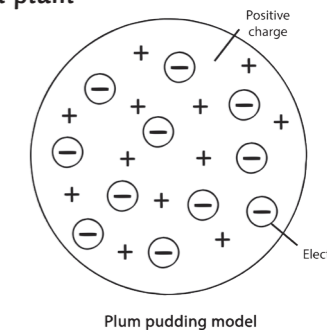
How are the groups arranged in the periodic table?  
**According to their properties.**

How can you tell that the alkali metals are very reactive?  
**According to their properties.**

How can you tell the noble gases are unreactive?  
**Full shell of outer electrons.**

Describe the plum pudding model of the atom.  
Draw a diagram.

**A sphere of positive charge with electrons dotted about; looking like a plum pudding.**



Why did scientists believe this model?  
**Lack of experimental evidence.**

Describe what the alpha scattering experiment showed scientists.

**Most alpha particles go straight through, some are scattered, some rebound off the gold foil. This shows that the nucleus of an atom has a very small radius. Most of the mass is concentrated in the nucleus.**

Niels Bohr discovered that  
**electrons orbit the nucleus in shells.**

Why did Mendeleev leave gaps in the periodic table?  
**He knew that the elements existed but they hadn't been found, based on their mass.**

What happened to some of the gaps he left?  
**They have been filled. Scientists have found some of the elements.**

The **transition** elements are a group of metals with similar properties which are different to the metals in group 1. Shade in the transition metals on the periodic table below.

Name three common transition metals.

**Students may name some of the following: chromium, manganese, iron, cobalt, nickel, copper or any other transition metal.**

Complete the sentences below to describe the properties of transition metals.

They form **coloured** compounds when reacting. They are **conductors** of heat and electricity. They are malleable. They have **high** densities.

Compared to the alkali metals, they are **less** reactive.

Complete the table to show the ions and colours formed by iron compounds. iron (III) oxide

Compound Name	Ion	Colour
iron (II) hydroxide	$\text{Fe}^{2+}$	pale green
iron (III) hydroxide	$\text{Fe}^{3+}$	orange-brown
iron (III) oxide	$\text{Fe}^{3+}$	red-brown

